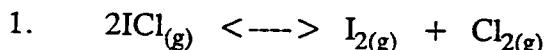


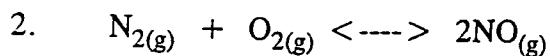
Equilibrium #4

Key.

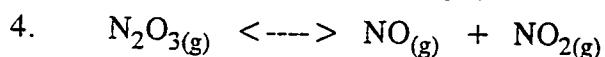
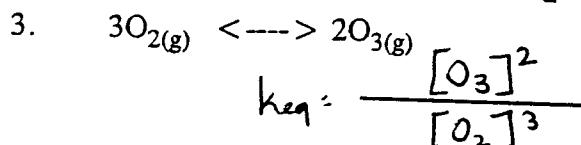
Write the equilibrium constant expression for each equilibrium.



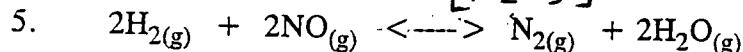
$$K_{\text{eq}} = \frac{[\text{I}_2][\text{Cl}_2]}{[\text{ICl}]^2}$$



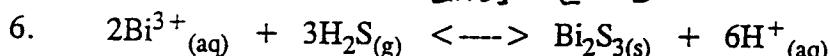
$$K_{\text{eq}} = \frac{[\text{NO}]^2}{[\text{N}_2][\text{O}_2]}$$



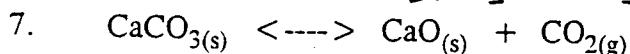
$$K_{\text{eq}} = \frac{[\text{NO}][\text{NO}_2]}{[\text{N}_2\text{O}_3]}$$



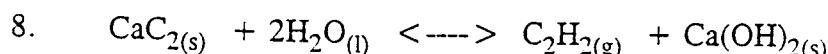
$$K_{\text{eq}} = \frac{[\text{N}_2][\text{H}_2\text{O}]^2}{[\text{H}_2]^2 [\text{NO}]^2}$$



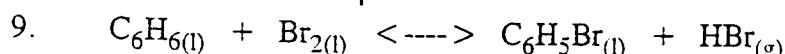
$$K_{\text{eq}} = \frac{[\text{H}^+]^6}{[\text{Bi}^{3+}]^2 [\text{H}_2\text{S}]^3}$$



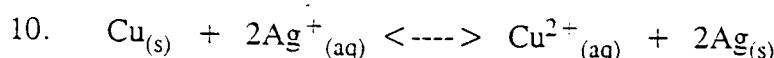
$$K_{\text{eq}} = [\text{CO}_2]$$



$$K_{\text{eq}} = [\text{C}_2\text{H}_2]$$



$$K_{\text{eq}} = \frac{[\text{C}_6\text{H}_5\text{Br}][\text{HBr}]}{[\text{C}_6\text{H}_6][\text{Br}_2]}$$



$$K_{\text{eq}} = \frac{[\text{Cu}^{2+}]}{[\text{Ag}^+]^2}$$

Key

