

KEY

Chemistry 11

Vocabulary List

ATOMIC MASS UNIT

- a unit for expressing masses in atoms, equal to $\frac{1}{12}$ of the mass of carbon 12. (p. 96)

ATOMIC MASS

- quantity of matter in an atom (p. 96)

MOLECULAR MASS

- The mass measurement that can be used only for covalent compounds.

FORMULA MASS

- The mass measurement that can be used for both covalent & ionic compounds.

ISOTOPE

= atoms of the same element that have different numbers of neutrons (p. 237)

MOLE

- an amount of matter that contains 6.02×10^{23} particles (p. 98)

MOLAR MASS

The mass (in grams) of 1 mole of atoms or molecules (p. 105)

AVOGADRO'S NUMBER

(N_A) (6.02×10^{23}) = a number used to express the relative mass of 1 mole of atoms (p. 98)

AVOGADRO'S HYPOTHESIS

- principal stating that equal volumes of different gases at the same temp. & pressure have equal numbers of particles (p. 97)

STP

- "standard temperature and pressure" (p. 182)
(an acronym) (101.3 kPa)
 (273 K)

MOLAR VOLUME

22.4 L at STP

(The volume of 1 mole of a gas at STP).