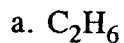
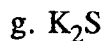


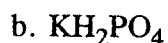
Calculate the molecular or formula masses of the following compounds:



$$\begin{array}{r} 2 \times 12.0 = 24.0 \\ 6 \times 1.0 = 6.0 \\ \hline 30.0 \text{ u} \end{array}$$



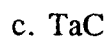
$$\begin{array}{r} 2 \times 39.1 = 78.2 \\ 32.1 \\ \hline 110.3 \text{ u} \end{array}$$



$$\begin{array}{r} 39.1 \\ 2 \times 1.0 = 2.0 \\ 31.0 \\ 4 \times 16.0 = 64.0 \\ \hline 136.1 \text{ u} \end{array}$$



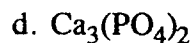
$$\begin{array}{r} 138.9 \\ 3 \times 35.5 = 106.5 \\ \hline 245.4 \text{ u} \end{array}$$



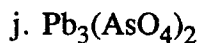
$$\begin{array}{r} 181.0 \\ 12.0 \\ \hline 193.0 \text{ u} \end{array}$$



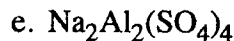
$$\begin{array}{r} 3 \times 12.0 = 36.0 \\ 6 \times 1.0 = 6.0 \\ 16.0 \\ \hline 58.0 \text{ u} \end{array}$$



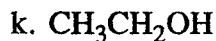
$$\begin{array}{r} 3 \times 40.1 = 120.3 \\ 2 \times 31.0 = 62.0 \\ 8 \times 16.0 = 128.0 \\ \hline 310.3 \text{ u} \end{array}$$



$$\begin{array}{r} 3 \times 207.2 = 621.6 \\ 2 \times 74.9 = 149.8 \\ 8 \times 16.0 = 128.0 \\ \hline 899.4 \text{ u} \end{array}$$



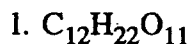
$$\begin{array}{r} 2 \times 23.0 = 46.0 \\ 2 \times 27.0 = 54.0 \\ 4 \times 32.1 = 128.4 \\ 16 \times 16.0 = 256.0 \\ \hline 484.4 \text{ u} \end{array}$$



$$\begin{array}{r} 2 \times 12.0 = 24.0 \\ 6 \times 1.0 = 6.0 \\ 16.0 \\ \hline 46.0 \text{ u} \end{array}$$



$$\begin{array}{r} 204.4 \\ 51.0 \\ 3 \times 16.0 = 48.0 \\ \hline 266.4 \text{ u} \end{array}$$



$$\begin{array}{r} 12 \times 12.0 = 144.0 \\ 22 \times 1.0 = 22.0 \\ 11 \times 16.0 = 176.0 \\ \hline 342.0 \text{ u} \end{array}$$